



Solvent Effects in MOLCAS

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1.- Effects of solvation on the potential energy surfaces

Solvatochromic effect: $\Delta E_{solvated} = \Delta E_{gas} + \Delta G^*_{excited} - \Delta G_{ground}$

2.- Reaction Field Methods: cavity + dielectric continuum

Spherical cavities: Kirkwood model Adapted cavities: Polarized Continuum Model (PCM)

3.- More advanced models:

QM/MM, etc



Solvent Effects

CASSCF/CASPT2, B. Proppe et al., J. Phys. Chem. A. 104, 1608 (2000)



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3.- Further models:

QM/MM, etc





ε=80

Methods that surround the quantum problem with electric fields, charges, dipoles or polarizabilities. *Implicit models: Reaction Field methods,PCM, etc*

- The environment is simulated by a polarizable continuum out of the cavity
- The continuum can be a dielectric, distributed polarizabilities, etc.
- The molecule reacts to the surroundings and create a reaction field which mimics the solvent effect.



- Energetic collapse: if large basis are used the charge can be placed outside the cavity boundaries, and the energy may collapse
- In general the result depends on the cavity size and they are basically qualitative
- Improvements: supermolecule approach (explicit solvation shell)









Onsager-Kirkwood Model: $H = H_0 + H_{RF} + H_{WF}$

• Spherical cavity (r = a) sorrounded by a dielectric continuum

$$E = E^{0} - \frac{1}{2} \sum_{l,m} M_{l}^{m} F_{l}^{m}, \quad F_{l}^{m} = c_{l} M_{l}^{m}$$
$$c_{l}(\epsilon) = -\frac{l!(l+1)(\epsilon-1)}{(l+1)\epsilon+l} \frac{1}{a^{2l+1}}$$



• Penalty gaussian functions (well potential) to avoid charge outside the cavity

$$E_{\text{ex}} = \iiint \rho(r, \theta, \phi) f(r) \, d\phi \, d\theta \, dr. \qquad f(r) = \sum_{i} \beta_{i} \exp(\alpha_{i} (r - R_{i})^{2})$$





Polarizable Continuum Model (PCM):

- Cavities as envelope of spheres centered on solute atoms or atomic groups
 - Pauling radii
 - UATM (united atom topological model) radii (default)
 - input sphere radii



 E_{int}

• Solvation charges placed in the middle of small tiles (*tesserae*) drawn on the surface; the number of solvation charges can be gauged by changing the average area of tesserae (AAre in **SEWARD**).



Solvent Effects for Excited States Nonequilibrium



Total reaction field of the ground state:

 $H_{\mathrm{RF,\,init}} = c_l(\boldsymbol{\epsilon}_0) M_{l,\,\mathrm{init}}^m$

Total reaction field of the excited state:

$$H_{\rm RF} = H_{\rm RF,\,iner} + H_{\rm RF,\,op}$$

$$E_{\text{sol}}^{\text{neq}} = \sum_{lm} F_{l, \text{op, fin}}^m M_{l, \text{fin}}^m + \sum_{lm} F_{l, \text{iner, init}}^m [2M_{l, \text{fin}}^m - M_{l, \text{init}}^m]$$

• Inertial or low response: reaction field of the fixed nuclei after excitation or emission

Colored from the ground (initial) state reaction field (it does not change)

• Optical or fast response: reaction field of the changed electronic charge after excitation or emission

Recomputed with the excited (final) state density using the optical dielectric ($\varepsilon = n^2$, n refraction index)





Onsager-Kirkwood Model in MOLCAS (ground states):

• SEWARD: definition of the cavity



• SEWARD: definition of the well potential





RFPert keyword. Reaction field computed by SCF or RASSCF read from RUNFILE added perturbatively





PCM Model in MOLCAS (ground and excited states):









Solvent Effects



Other methods to simulate the solvent



Methods that combine quantum and classical mechanics: Explicit models: Quantum Mechanics / Molecular Mechanics (QM/MM)

 $H_{complete} = H_{QM} + H_{MM} + H_{QM/MM}$

ONIOM (in Gaussian): is divided in shells, internal and more accurate QM, and external charges or dielectric with MM

AM1/OPLS, QM/RISM, XSOL, etc..



The results will depend on the the accuracy of the force fields parametrized for MM, of the quality of the QM calculation, of the coupling of the shells, and of the design of the separation.

Molcas: new ESPF code: QM + (Electrostatic potential, field, field deriv.)

Coupled to Tinker or Amber

Methods that use exclusively quantum mechanics: (QM/MD) Car-Parrinello (1985)







Basic quantum-chemical solvation effects

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